Test Jot Series

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Student Name:			**Class:**	9th
Paper Type:	**Test Series**		**Subject:**	**Chemistry (SNC)**
Paper Time:	**45**		**Maximum Marks:**	31

Q1. Choose the correct option. (7X1=7)

1. The colour of hydrated copper sulphate solid is:

(A) Blue	**(B)** Black	**(C)** Pink	**(D)** White

2. Such reaction which continue in both direction are called?

(A) Irreversible	**(B)** Reversible	**(C)** Non-reactive	**(D)** Dynamic

3. The reaction in which the products do not recombine to form reactants are called?

(A) Irreversible reactions	**(B)** Reversible reactions	**(C)** Decomposition reactions	**(D)** Addition reaction

4. What will happen if the rates of forward and reverse reactions are very high?

(A) The equilibrium point will reach very soon	**(B)** The equilibrium point will reach will reach very late	**(C)** The reaction will not attain the sate of dynamic equilibrium	**(D)** The reaction will be practically irreversible

5. What is the relationship between the equilibrium constant (k) and the concentrations of reactants and products at dynamic equilibrium?

(A) k= (reactions) (products)	**(B)** k= (products) (reactions)	**(C)** k= (reactions) (products)	**(D)** k= (products) (reactants)

6. Industrially ammonia is produced by which process?

(A) Haber process	**(B)** Solvay process	**(C)** Hydrogenation	**(D)** Halogenation

7. Nitrogen and hydrogen were reacted together to make ammonia? $N_2+3H_2 \rightleftharpoons 2NH_3$ kc=2.86 mol^{-2} dm^6

(A) NH_3 only	**(B)** N_2, H_2 and NH_3	**(C)** N_2 and H_2 only	**(D)** H_2 only

Q2. Write down short answers of following questions. (7X2=14)

[i] Why are combustion reactions generally irreversible? (Constructed Response Questions)	**[ii]** What are reversible reaction?
	[iii] What is copper sulphate penta hydrate?
[iv] How is dynamic equilibrium different form the static equilibrium?	**[v]** How can you decrease the time to attain the position of equilibrium in a reversible reaction?
[vi] How can you drive the reversible reaction at equilibrium? (a) in the forward direction (b) in the backward direction	**[vii]** What is the effect of increasing pressure on the following reaction at equilibrium? $N_{2_g} + O_{2_g} \rightleftharpoons 2NO_{(g)}$

Q3. Write detailed answers of the following questions.(Answer any **2**) (2X5=10)

1. Explain how this reaction is reversible? $CaCO_3 \rightleftharpoons CaO+CO_2$
2. How can a reversible reaction be forced to go completion?
3. Explain five conditions that can disturb the chemical equilibrium state?