

Test Jot Series

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Student Name:			Class:	9th
Paper Type:	Test Series		Subject:	Chemistry (SNC)
Paper Time:	45		Maximum Marks:	30

Q1. Choose the correct option. (6X1=6)

1. Which compound has same molecular and empirical formula?

(A) H_2O_2	(B) C_6H_6	(C) $C_6H_{12}O_6$	(D) H_2O
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2. Formula of Magnesium nitride:

(A) Mg_3N	(B) Mg_2N_3	(C) MgN_2	(D) Mg_2N_2
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3. What is the chemical formula for carbon dioxide?

(A) CO	(B) CO_2	(C) C_2O	(D) C_2O_2
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4. What is the chemical formula for calcium carbonate?

(A) $CaCO_2$	(B) $CaCO_3$	(C) Ca_2CO_3	(D) Ca_3CO_2
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5. What is the term for the substances formed in a chemical reaction?

(A) Reactants	(B) Catalysts	(C) Products	(D) Intermediates
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6. 40 g of H_3PO_4 contains number of moles?

(A) 0.58g	(B) 4.8g	(C) 0.408g	(D) 5.8g
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Q2. Write down short answers of following questions. (7X2=14)

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| [i] What is the relationship between molecular and empirical formula? | [ii] What is the molecular formula of acetic acid? |
| [iv] How can you calculate the masses of the products formed in a reversible reaction? (Constructed Response Question) | [v] Difference between reactants and products? |
| [vii] What is balanced chemical equation? | [vi] What is the number of moles of O_2 required to burn 2 moles of methane. $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$? |
| [iii] Calculate the molar mass of $MgCO_3$? | |

Q3. Write detailed answers of the following questions. (Answer any 2) (2X5=10)

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| 1. A compound has the following formula $CH_3-CH_2-CH_2-COOH$. determine the empirical formula? |
| 2. When 8.657g of a compound were converted into elements, it gave. 5.217g of carbon 0.962g of hydrogen and 2.478g of oxygen. Calculate the percentage of each element present in this compound. (Constructed Response Questions) |
| 3. Explain the concepts of Avogadro's number and mole. |