

Test Jot Series

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Student Name:			Class:	9th
Paper Type:	Test Series		Subject:	Chemistry (SNC)
Paper Time:	45		Maximum Marks:	30

Q1. Choose the correct option. (6X1=6)

1. A bond formed between two non metals is expected to be:

(A) Covalent	(B) Ionic	(C) Metallic	(D) All of these
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2. Metals have low values of:

(A) Electronegativity	(B) Electron affinity	(C) Ionization energy	(D) None of these
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3. What is the primary force responsible for holding metal atoms together in a metallic bond:

(A) Electrostatic attraction	(B) Ionic bonding	(C) Covalent bonding	(D) Vander waals
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4. Example of ionic compound is:

(A) CH ₄	(B) H ₂ SO ₄	(C) C ₆ H ₁₂ O ₆	(D) NaCl
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5. Conductivity of ionic compounds in due to free:

(A) Electrons	(B) Protons	(C) Neutrons	(D) None
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6. Which is covalent compound:

(A) CH ₄	(B) H ₂ SO ₄	(C) C ₆ H ₁₂ O ₆	(D) All of these
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Q2. Write down short answers of following questions. (7X2=14)

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|---|---|
| [i] Why do Atoms form chemical bond? | [ii] Write down dot and cross formula of HNO ₃ . |
| [iii] How do metals conduct heat? (Constructed Response Question) | [iv] Write down two properties of metals? |
| [vi] Define covalent compound? Give example. | [v] Name two most electronegative element? |
| | [vii] Why ice floats on water? |

Q3. Write detailed answers of the following questions.(Answer any 2) (2X5=10)

- Write a brief note on covalent bond? Explain its types?
- Explain the properties of metals keeping in view the nature of metallic bond.
- What is metallic bond? Explain diagram?