

Test Jot Series

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Student Name:			Class:	9th
Paper Type:	Test Series		Subject:	Chemistry (SNC)
Paper Time:	45		Maximum Marks:	30

Q1. Choose the correct option. (6X1=6)

1. How many electrons can be accommodated at the most in the third shell of the elements?

(A) 8	(B) 18	(C) 10	(D) 32
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2. Proton carry charge:

(A) Positive	(B) Negative	(C) Neutral	(D) All of these
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3. Which particle has no charge:

(A) Proton	(B) Neutron	(C) Electron	(D) All of these
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4. If oxygen has 8 protons in it its Atomic Number will be:

(A) 8	(B) 16	(C) 18	(D) 10
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5. Formula of atomic number:

(A) $Z=A-n$	(B) $A=Z+n$	(C) $A=Z-n$	(D) $Z=A+n$
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6. Which isotope of hydrogen have zero neutron:

(A) Protium	(B) Deuterium	(C) Tritium	(D) None of these
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Q2. Write down short answers of following questions. (7X2=14)

[i] Why does the energy of electron increase as we move [ii] What is the structure of an atom?

from first shell to second shell? (Constructed Response Question) [iii] Differentiate between shell and sub shell with examples of each.

[iv] Who discovered an electron and a neutron?

[v] How many neutrons are present in ${}_{83}^{210}\text{Bi}$

[vi] Difference between atomic number and atomic mass?

[vii] How can an atom absorb and involve energy?

Q3. Write detailed answers of the following questions.(Answer any 2) (2X5=10)

1. Calculate the number of neutrons, protons and electrons in following atoms, ${}_{78}^{195}\text{Pt}$, ${}_{25}^{55}\text{Mn}$, ${}_{53}^{127}\text{I}$

2. What is radioactivity? Explain any three applications of radioactive isotopes.

3. Calculate the relative atomic mass of lead (Pb). isotopic abundances of isotopes are 2.0, 24.0, 22.0 52.0 respectively?